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題:販売規制G法規

R:559 P. 04/18

Document 1

アロニックス®
Mシリーズ

アクリル系特殊モノマー・オリゴマー

東亞合成化学



08年10月03日(金) 14時54分 案件番号:0-0362039480

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■アロニックスとは

アロニックスは、東亜合成化学が開発したアクリル系特殊モノマー・オリゴマーの商品名です。

現在市販されているアロニックスには、特殊アクリレート、ウレタンアクリレート及びポリエステルアクリレートがあります。

特殊アクリレート

低粘度であり、皮膚刺激性の低いものが多く、光硬化性が優れているなどの特徴を持っています。従って特殊アクリレートは、アクリル系オリゴマーの粘度を低下させたり、アクリル系オリゴマーの接着性、耐熱性、硬度及び硬化性を向上させるための反応性希釈剤として有効です。

ウレタンアクリレート

ウレタン結合を主鎖に持ち、強靭性に優れた塗膜が得られ、各種素材との接着性が優れるなどの特徴を持っています。

ポリエステルアクリレート

エステル結合を主鎖とし、アクリル系不飽和結合を分子内に1~数個持っており、他のアクリル系オリゴマーと比べて、比較的低粘度で液体のものが容易に得られ、他のポリマーやオリゴマーと相溶性が優れるなどの特徴を持っています。

アロニックスは、これらの特徴を生かすことにより、各種の硬化手段を用いることができます。

また、硬質から軟質タイプまで用途・機能に合わせ幅広い分子設計が可能です。

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アロニックス

■アロニックスの種類

アロニックスは、樹脂の構造、アクリロイル基の数により下表のように分類されます。

特殊アクリレート

アクリロイル基の数		アクリロイル基の数	アクリロイル基の数	アクリロイル基の数
単官能	M-101	フェノールEO変性(n=4)アクリレート		低粘度
	M-102	フェノールEO変性(n=4)アクリレート		低毒性
	M-111	メチルフェノールEO変性(n=3)アクリレート		良柔軟性
	M-113	メチルフェノールEO変性(n=4)アクリレート		
	M-117	メチルフェノールPO変性(n=5)アクリレート		
	M-120	2-エチル-2-キシルカルボキシルアクリレート	N-VP	
2官能	M-150	N-ビニル-2-ビロリドン	A-EPE4	
	M-210	ビスフェノールA EO変性(n=2)ジアクリレート		
	M-215	イソシアヌル酸EO変性ジアクリレート	TPGDA	低粘度
	M-220	トリプロピレンジグリコールジアクリレート	TEGDA	
	M-233	ベンタエリスリトールジアクリレートモノスチアレート	PEG #200	
	M-240	テトラステレンジグリコールジアクリレート(n=4)	PEG #400	
	M-245	ポリエチレンジグリコールジアクリレート(n=6)	PEG #600	
3官能	M-260	ポリエチレンジグリコールジアクリレート(n=14)	PPGDA	
	M-270	ポリプロピレンジグリコールジアクリレート(n=18)		
	M-305	ベンタエリスリトールジアクリレート	PETA	低粘度
	M-309	トリメチロールプロパンEO変性(n=1)トリアクリレート	TMPTA	
	M-310	トリメチロールプロパンEO変性(n=1)トリアクリレート		
	M-315	イソシアヌル酸EO変性トリアクリレート		
官能	M-320	トリメチロールプロパンEO変性(n=2)トリアクリレート		高柔軟性
	M-350	トリメチロールプロパンEO変性(n=1)トリアクリレート		
	M-360	トリメチロールプロパンEO変性(n=2)トリアクリレート		
	M-400	シベンタエリスリトールベンタ及バキサアクリレート	DPHA	高柔軟性
多官能	M-450	ベンタエリスリトールテトラアクリレート		高硬化性
				高硬度



—ウレタンアクリレート—

官能基数	M-1100 M-1200 M-1210 M-1310 M-1600	黄変型 無黄変型 無黄変型 黄変型 無黄変型	中硬質タイプ 中硬質タイプ 低ヤング率 軟質タイプ 速硬化・中硬質タイプ
2官能			

特殊アクリレート

官能基数	M-5000 シリーズ	-COOH, -OH 基 含有モノマー
單官能		

—ポリエステルアクリレート—

官能基数	M-6000 シリーズ	低毒性, 低粘度
2官能	M-7000 シリーズ	低毒性, 高光沢, 高硬化性
多官能	M-8000 シリーズ M-9000 シリーズ	低毒性, 高硬度, 高硬化性 耐熱性, 高硬度, 高硬化性

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归属：東亞合夥物業G名古屋

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Document 2

New Values Of the Solubility Parameters From Vapor Pressure Data

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The solubility parameters of a broad spectrum of solvents and chemicals are calculated from vapor pressure data using an expression derived from the relationship of Hugginsmacher. In the case of high boiling liquids, the available vapor pressure data are found to be unreliable when extrapolated to room temperature. A new method of calculation is proposed. A structure correlation is made using the method of Small and new values of the molar cohesion constants are developed. The problem of associations of certain molecular species is discussed and the concept of chameleonic character introduced as a qualitative explanation.

INTRODUCTION

Although a few coatings chemists have long recognized the value of the contributions of Hildebrand, Flory, Scott, Huggins, and others, in the areas of polymer solubility and compatibility—only in recent years have the developed theories been applied to solve practical problems encountered in the coatings industry. Perhaps the reason for the long span of time between theory and application—nearly 60 years since the initial work by van Laar—has been the inability of the theory to adequately predict the behavior of the complex mixtures required for practical vehicle formulations. However, as the coatings industry becomes more sophisticated, the principles set forth by these and other workers are being increasingly adopted.

The mathematical description of the solution process as first proposed by van Laar and later developed by Hildebrand¹ and Scatchard² is centered around a concept introduced, and termed the "solubility parameter," by Hildebrand.

* Coatings Materials, South Charleston, W. Va.

The classic compilation of data by Burrell³ in middle of the last decade has served as a basis for extending the usefulness of the concept by more recent workers including Lieberman,⁴ Fraunitz,⁵ Cordon,⁶ and Crowley *et al.*⁷ However, since the work of Burrell there has been no concerted effort to refine and expand this basic information.

The initial purpose of this work is not to seek explanations for notorious exceptions to the theory but rather to re-examine the source of the data and refine and expand it to a much broader working base.

DISCUSSION

Calculation of Solubility Parameter From Vapor Pressure Data

The solubility parameter is defined as:

$$\delta = \left(\frac{\Delta E}{V} \right)^{1/2}$$

where δ is the solubility parameter, ΔE the internal energy, and V the molar volume. Since $\Delta E = \Delta H - \Delta V$ the solubility parameter becomes:

$$\delta = \left(\frac{\Delta H - \Delta V}{V} \right)^{1/2}$$

In the final analysis the calculation of the solubility parameter revolves around obtaining the value of the heat of vaporization. Burrell used an empirical equation developed by Hildebrand for hydrocarbons and assumed the vapors behaved ideally. The parameter was calculated by Burrell and then corrected according to several empirical rules he developed.

It was decided to recalculate values for the solubility parameters from vapor pressure data and to extend the list as much as possible.

A search of the literature reveals that one of the

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more accurate expressions for calculating the heat of vaporization at a given temperature from available vapor pressure data is given by the Haggengmacher's equations, (3) and (4):

$$P(V_c - V_l) = \frac{RT}{M} \sqrt{1 - \frac{P_{c}T^2}{P_{c}T^2}} \quad (3)$$

$$\Delta H = \frac{dp}{dt} \frac{RT^2}{M} \sqrt{1 - \frac{P_{c}T^2}{P_{c}T^2}} \quad (4)$$

where V_c = the specific volume of the gas phase
 V_l = the specific volume of the liquid phase
 M = molecular weight
 R = gas constant
 t = temperature
 T_c = critical temperature
 P_c = critical pressure
 P_{cT} = critical pressure
 ΔH = heat of vaporization

Using equations (3) and (4) and vapor pressure in the form of the Antoine equation (5):

$$\log P = \frac{-B}{t+C} + A \quad (5)$$

where P = pressure in mm Hg; t = temperature in °C; and A , B , C are constants, it is possible to derive an expression for the solubility parameter as shown in equation (6):

$$\chi = \left[\frac{RT}{M} \sqrt{1 - \frac{P_{c}T^2}{P_{c}T^2}} \frac{2.303 \Delta H^2}{(t+C-273.15)^2} - 1 \right]^{1/2} \quad (6)$$

Consequently, the following data are necessary to calculate the solubility parameter of a pure liquid at any desired temperature: Antoine constants A , B , and C , critical temperature (T_c), critical pressure (P_c), density (ρ), temperature (t) in degrees abs., and molecular weight (M).

Generally it is desirable to obtain the value of the solubility parameter at 25°C. In certain cases this temperature is beyond the range of the usual Antoine expression. An examination of the Antoine equation indicates the source of the error, illustrated by the somewhat exaggerated diagram shown in Figure 1. At pressures between P_1 and P_2 the Antoine equation describes the vapor pressure-temperature relationship quite well. However, at lower and higher temperatures the deviation actually becomes intolerable and usually new sets of Antoine constants are obtained. Unfortunately, these new Antoine constants are not always available.

To circumvent this problem, an alternate means of reliably estimating the heat of vaporization at room temperature, from data at different (usually higher) values, is necessary. It has been found empirically that at pressures below atmospheric pressure the latent heat of vaporization follows the relationship:

$$\Delta H = \Delta H^{\circ} e^{-\chi t} \quad (7)$$

In logarithmic form this becomes:

$$\log \Delta H = -\frac{m}{2.303} t + \log \Delta H^{\circ} \quad (8)$$

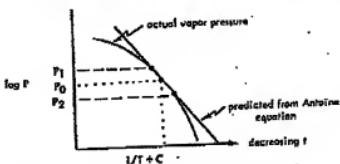


Figure 1—Illustration of fit of Antoine equation to experimental data (exaggerated for visual interpretation)

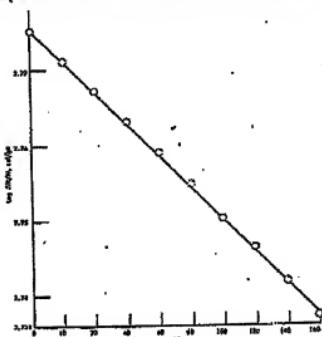


Figure 2—Effect of temperature on $\log (\Delta H/M)$ of water

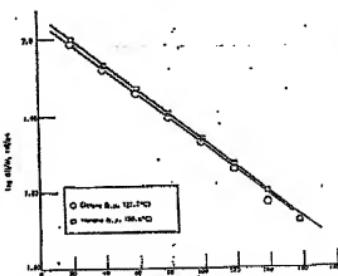
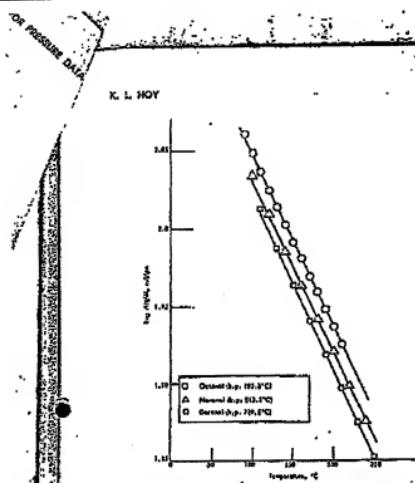
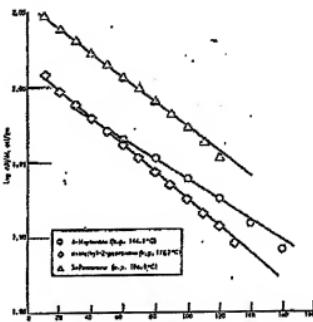
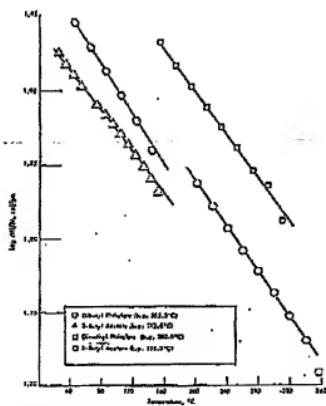


Figure 3—Effect of temperature on $\log (\Delta H/M)$ of normal hydrocarbons

Figure 4—Effect of temperature on $\log (\Delta H/M)$ of alcoholsFigure 6—Effect of temperature on $\log (\Delta H/M)$ of ketonesFigure 5—Effect of temperature on $\log (\Delta H/M)$ of esters

where ΔH° is the heat of vaporization at some standard temperature and m is a constant. A test of equation (8) is shown in Figures 2-6 for several different classes of compounds. As can be seen, the relationship seems to be quite general.

Using this relationship it is possible to estimate the heat of vaporization at 25°C by calculating the heat of vaporization in the temperature range in which the Antoine constants are valid and fitting these values into equation (8) to determine the slope (m), and (ΔH°). With these two constants the heat of vaporization at 25°C can be estimated. Table I shows the results of such an extrapolation for a variety of compounds. Based on these results it is believed that the maximum error is above ± 45 or translated to a solubility parameter, approximately ± 25 .

The task of carrying out the calculations is relegated to a high speed digital computer. The data are tabulated in three forms: (1) an alphabetical listing, (2) a listing in order of increasing solubility parameters, and (3) in order of increasing boiling points. In addition to solubility parameters and boiling points, data on molecular weight, vapor pressure at 25°C and density in lbs./gallon at 25°C are included. Solubility parameter data for more than 650 compounds have been calculated and are shown in the tabulations. Some materials not normally considered solvents are included because of the usefulness of these data to chemists interested in finding reaction media for various processes. Moreover, it was desired to obtain as many different structural types as possible—for a systematic correlation of structure and solubility parameter.

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Table 1—Comparison of Estimated and Observed Values of Solubility Parameters

Index No.	Compound	H _{vap} (cal/mole) ^{1/2}	S ₀₅ (cal/cm ³) ^{1/2} /mole	δ^* (cal/cm ³) ^{1/2}	% Deviation in δ
.752	Acrylic acid	(Estimated)	16.308.40	1511.7	11.588
.192		(Observed)	16.597.41	1521.3	0.90
.922	Methyl butanediol	(Estimated)	15.902.83	1510.1	18.374
.714		(Observed)	16.989.57	1516.4	12.456
.327	Octane	(Estimated)	15.260.19	1516.5	7.818
		(Observed)	15.265.77	1512.5	1.67
.121	Dimethylbenzene	(Estimated)	15.570.99	1535.2	6.815
.763		(Observed)	15.221.19	1534.8	8.712
.361	p-Ethylstyrene	(Estimated)	12.707.65	1545.0	9.015
.707		(Observed)	12.421.65	1507.3	8.920
.031	Isomethylstyrene	(Estimated)	13.812.76	1519.4	10.020
.862		(Observed)	13.540.33	1504.6	9.922
.228	1,1-Dichloroethene	(Estimated)	13.555.43	1523.7	10.005
.725		(Observed)	13.522.23	1521.9	9.931

(1) Sum of Small's constants = $\frac{8.3 \times M_w}{4}$ (2) δ = Solubility parameter.

Structure Correlations

Small¹ developed a method for estimating the value of the solubility parameter from chemical structure through additive group and constitutive molar attraction constants.

One of the objectives of the present project is to re-examine these molar attraction constants over a broad spectrum of compounds using multiple regression analysis. This is readily accomplished with high speed digital computers. In the course of this investigation it became apparent, as was expected, that the only compounds which did not reliably yield to analysis were the acids, alcohols and other compounds which are capable of association. The carboxylic acids, for example, were predictable only to the extent of $\pm 18\%$. It occurred to the author that since the simple analysis was based on monomeric acids, not the dimeric form, if the structural features of the dimers were

taken into account these compounds might fulfill the requirements for a successful analysis.

If it is assumed that carboxylic acids exist as dimers, then the following factors must be incorporated in the calculation of the solubility parameter. Equation (1) for solubility parameter can be expressed as:

$$\delta = \left(\frac{E^*}{M} \right)^{1/2} \quad (1)$$

where δ = solubility parameter (cal/cm³)^{1/2}
 E^* = molar internal energy (cal/mole)
 M = molecular weight (gm/mole)
 ρ = density (gm/cm³)

Table 2—Correlation of Molar Volume Cohesion
Of Glycol Esters

Compound	Molar Volume Cohesion (cal/cm ³) ^{1/2} /mole		Error
	(Observed)	(Calculated)	
Acetic acid	1053.0	1014.5	5.5 %
Acrylic acid	1259.1	1301.1	2.8 %
Butyric acid	1346.3	1346.1	1.5 %
3-Ethylpropionic acid	2049.5	2027.5	0.6 %
3-Ethylacrylic acid	2056.7	2005.5	2.4 %
3-Ethylhexanoic acid	3197.8	2537.1	2.6 %
2-Ethyl-3-propyl acrylic acid	3615.6	2182.5	0.9 %
		2077.9	0.6 %
Isobutyric acid	1560.9	1473.9	6.7 %
Methacrylic acid	1581.1	1561.5	1.2 %
2-Methyl pentanoic acid	1753.9	1802.5	6.5 %
Propionic acid	1820.7	1292.5	3.2 %
Tetrahydrobenzoic acid	3082.0	3099.5	0.85 %
Valeric	1815.3	1811.9	19 %

Compound	Molar Volume Cohesion (cal/cm ³) ^{1/2} /mole		Error
	(Observed)	(Calculated)	
Benzyl Carbonyl	1565.8	1577.7	1.4 %
Propyl Carbonyl	1501.4	1514.2	0.98 %
Cellosolve solvent	1045.7	1048.4	0.67 %
2-Ethylbutyl Cellosolve	1555.1	1545.5	0.72 %
2-Ethylpropyl Cellosolve	1795.0	1810.0	0.85 %
3-Methyl Cellosolve	1522.2	1576.9	3.3 %
Heptyl Cellosolve	1629.3	1580.0	0.56 %
α -Methylbenzyl Cellosolve	1611.9	1574.5	2.6 %
Methyl Cellosolve	995.3	975.5	1.1 %
Propyl Cellosolve	1570.6	1570.6	0.00 %
Butyl Carbonyl	1441.0	1444.7	0.56 %
Carboxil solvent	1676.5	1687.9	0.67 %
	1418.3	1422.1	0.68 %
2-Ethylbutyl Carbonyl	1571.0	1574.5	1.0 %
2-Ethylpropyl Carbonyl	1599.1	2133.4	0.58 %
Heptyl Carbonyl	1590.5	1593.7	1.6 %
Methyl Carbonyl	1520.4	1286.2	2.1 %
Phenyl Carbonyl	1524.7	1518.4	0.38 %
Propyl Carbonyl	1536.9	1554.1	1.17 %

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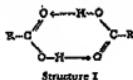
In the case of dimeric carboxylic acids, however, the actual molecular weight is two times the simple formula weight (M), and the solubility parameter becomes:

$$\delta = \left(\frac{\Delta E}{M} \right)^{1/2} \cdot \sqrt{\frac{E}{2}} \quad (10)$$

In a like manner the molar-volume cohesion (G) for dimeric carboxylic acids is given by equation (11):

$$G = \left(\Delta E \frac{M}{M} \right)^{1/2} \cdot \sqrt{\delta} \quad (11)$$

The value of the solubility parameter of acetic acid when calculated from vapor pressure data and considered as an unassociated molecule is 13.01. However, when considered as a dimer, its value is 9.18. It is apparent that the dimeric form will have quite different solubility characteristics from the monomeric form, and it is therefore understandable that acetic acid is soluble in such diverse solvents as water ($\delta = 23.5$) as well as heptane ($\delta = 7.5$). Treatment of a series of 14 carboxylic acids by multiple regression analysis using the dimeric Structure I has given satisfactory results, as listed in Table 2.



Another group of solvents which have evaded correlation by simple structure considerations are the glycol-ethers.* In these cases it is possible to postulate the



intramolecular-hydrogen bonded Structure II. Using this structure it is possible to correlate the molar volume cohesion by regression analysis. The results of these correlations are shown in Table 3.

In this analysis, as well as in the case of the acid, over 640 organic compounds were considered. The samples consisted of aliphatic hydrocarbons, aromatic hydrocarbons, ketones, esters, and ethers. The overall analysis is the most reliable to date and accounted for 98.20% of the residual sum of squares. The molar attraction constants generated are given in Table 4.

Interestingly, a cursory examination of the solubility characteristics of the glycol-ethers reveals that certain members of these series, like the auids, are capable of being dissolved in polar solvents (water) and non-polar solvents (heptane). See Table 5 for effects of these and other structures. It would appear

*Cetolene and Carbilex solvents, products of Union Carbide Corp.

Table 4—Molar Attraction Constants

Group	Molar Attraction (S) (cal $\text{cm}^{3/2}$ /mole)	Group	Molar Attraction (S) (cal $\text{cm}^{3/2}$ /mole)
CH_3	146.5	S^-	200.42
CH_2	131.5	Cl^-	243.67
$>\text{CH}$	85.99	Cl primary	205.06
$\text{C}=\text{C}$	32.03	Cl secondary	209.27
$\text{CH}_2=\text{O}$	128.54	Cl aromatic	161.0
$\text{CH}_2=\text{N}$	121.53	Br	297.85
$>\text{C}=\text{O}$	84.51	Br aromatic	205.60
$\text{CH}_2=\text{ aromatic}$	112.12	F	41.53
C_6H_5	98.12	Structure feature	
$\text{C}_6\text{H}_5\text{O}$	114.98	Conjugation	81.56
$\text{C}_6\text{H}_5\text{OCH}_3$	176.53	C_6H_5	-7.13
COO	282.58	Twenz	15.50
$>\text{C}=\text{O}$	226.64	4 Membered ring	77.76
CHO	267.29	5 Membered ring	20.39
$(\text{CO})_2\text{O}$	223.84	6 Membered ring	23.44
OH^-	223.47	Ortho substitution	5.5
H_2O acidic dimer	170.59	Meta substitution	40.33
OH aromatic	225.53	Para substitution	22.5
NH_3^+	180.05	Base Value	155.1
NH_2	61.08		
N^-	254.26		
C_6N	328.06		

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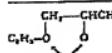
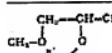
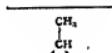
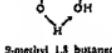
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SOLUBILITY
[CONT'D.]

PRESSURE DATA

C. L. HOY

Table 5—Effect of Structure on Solubility
In Water and Hydrocarbon Solvent (Cont'd.)

Structure	Solubility In	
	H ₂ O	Resins
	∞	∞
	∞	∞
	∞	∞
	∞	∞

that if the energy requirements can be fulfilled either by simple dimerization or intramolecular association, then that part of the molecule tends to adopt the character of the surrounding environment; i.e., in polar solvents the materials are capable of interacting as polar solvents, while in non-polar solvents the polar interactions are self-contained and the materials tend to behave in a non-polar manner. It is proposed that this ability to assume the character of the surrounding environment be termed "chameleonic" after the reptile which is able to assume the color of his background.

The chameleonic principle has wide ramifications when applied to practical problems; for example, the puzzling compatibility of the vinyl resin VAGH,¹ as compared to vinyl resin VYHH.² VYHH is a vinyl chloride-vinyl acetate copolymer and VAGH is a back hydrolyzed version of VYHH. Yet VAGH has a broad

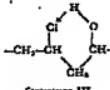
¹ Backlined vinyl resin, VAGH, product of Union Carbide Corp.

² Backlined vinyl resin, VYHH.



DR. C. L. HOY received the B.S. Degree from the University of Wyoming in 1930, and the Ph.D. Degree from the University of North Carolina in 1935. In 1935 he joined Union Carbide Corp. and is currently in the Color and Coatings Research and Development Department. He has worked on vinyl polymers, solvents, plasticizers, urethanes, epoxides, and alkyls, with special emphasis on dispersion. At present his group is responsible for fundamental work in dispersion systems.

spectrum of compatibility with polar and relatively non-polar solvents and resins in contrast to VYHH. Further inclusion of other hydroxy containing monomers in the polymer do not produce the dramatic effect of the vinyl alcohol moiety. However, if the vinyl chloride-vinyl alcohol entity operates as a chameleonic Structure III then it would be expected that VAGH would have a unique broad spectrum compatibility and solubility which can not be duplicated by the use of other monomers such as hydroxy ethyl acrylate.



Structure III

This explains why certain glycol-ether^{**} solvents are useful as coupling solvents for stabilizing relatively non-polar resins in water. The use of glycol-ethers^{††} to prevent blushing and cratering of non-polar resins during drying under humid conditions can also be explained by the chameleonic principle.

SUMMARY

Over 680 values of the solubility parameter have been calculated from vapor pressure data and presented in tabular form. From the data obtained it has been possible to expand the usefulness of Smill's Molar Cohesion constants. The chameleonic nature of associated molecules is discussed as a means of understanding certain observed physical phenomena, particularly puzzling solubility characteristics, of solvents which have evaded simple structure correlations.

ACKNOWLEDGMENT

The authors would like to acknowledge the invaluable work of Dr. R. A. Martin without whose service the mathematical analysis via computer would not have been possible. *

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^{**} Selective Carbonyl, Odeolive, and Propylene glycols, products of Union Carbide Corp.

^{††} Gallyol and Carbital solvents.

Document 3

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(51) [564]

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ポリマーのSP値の予測—訂正と補足

井 本 一 稔

先に本誌に接着剤のことを書き、ポリマーのSP値を予測する計算法のことを簡単に記した。そのさい不正確にも既述の名を書かれた(文獻番号24b)Small, I. Appl. Polym. Sci. は J. Appl. Chem. であった。そのこともあり、あらためてポリマーのSP値を予測するSmallとHayとEdgarの報告をもう少し詳しく書いておきたいと考える。

1. Smallの場合

1953年に出了報告で後は溶液の熱力学から論じはじめる。きっちり詰められた5ページを超える論文は基礎的で時間をかけて勉強するに足る。しかし、ここはすべて省略し、SP値の予測の所だけに限る。Smallのその式は簡単式(1)である。「ポリマーのくり返し単位についてである。

$$SP = \delta = \frac{\sum F}{V} \quad (1)$$

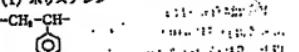
F というのは“molar-attraction constant”と彼が呼ぶもので、表の値である。Vはモル容積($\text{cm}^3 \text{ mol}^{-1}$)である。また F の単位は $(\text{cal cm}^3)^{1/2} \text{ mol}^{-1}$ である。

表1 $F(\text{cal cm}^3)^{1/2} \text{ mol}^{-1}$ の数値(Small)

CH_2	214	H_2	80-100
CH_2	133	O	70
CH_2	28	PCl_3	275
CH_2	(1)-93	$\text{C}-\text{O}-$	310 (エステル)
CH_2	190	CN	410
CH_2	111	Cl	270-260
CH_2	19	$\text{O}-\text{Cl}$	260
CH_2	285	$\text{C}-\text{Cl}$	250
CH_2	222	Br	340
CH_2	735	I	425
C_2H_5		CP	150
C_2H_5	(o, m, p)	558	
C_2H_5		CP_2	274
ナフチル	1145	CP_3	
5風呂	105-115	S	225
6風呂	99-105	SH	315
共役系	20-30	ONO_2	440
		NO_2	440

計算例をあげた方が早かりる

例(1) ポリстиレン

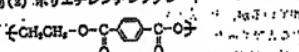

 $\text{C}_6\text{H}_6, M=104$
 $d=1.05$
実測 $SP=80$

$$\sum F = 133(\text{CH}_2) + 28(\text{CH}) + 735(\text{C}_6\text{H}_6) = 896$$

$$V=104/105=99$$

$$\text{よって} \quad \text{計算 } SP = 896/99 = 9.05$$

例(2) ポリエチレンテレフタレート



$$\sum F = 2 \times 133(\text{CH}_2) + 2 \times 310(\text{COO}) + 658(\text{C}_6\text{H}_4)$$

$$d=138, \text{実測 } SP=107, 1 \text{に対する計算は下記のとおりである。}$$

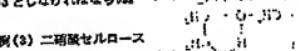
$$\sum F = 2 \times 133(\text{CH}_2) + 2 \times 310(\text{COO}) + 658(\text{C}_6\text{H}_4)$$

$$d=138, \text{実測 } SP=107, 1 \text{に対する計算は下記のとおりである。}$$

$$V=192/138=1.3821$$

よって計算 $SP=110$ Small は $SP=107$ と計算しているが、それには $d=133$ としなければならぬ。

例(3) 二硝酸セルロース



$$\sum F = 2 \times 133(\text{CH}_2) + 2 \times 310(\text{COO}) + 658(\text{C}_6\text{H}_4)$$

$$d=224, \text{実測 } SP=105.1$$

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ポリマーのsp値の予測一訂正と補足

表2 Hoyの法の例(680種のうちから)

分子量	沸点 (760 mm)	臨界圧 (25°C, MM)	SP*	比重 (25°C, lb./Gal.)**
ジエチレンジアコール	106.12	245.7	0.00	14.24 9.286
トルレン・ジイソシアナート	174.17	238.3	0.00	11.50 10.116
アセトニトリル	41.05	81.1	85.33	12.11 6.478
MMA	99.12	102.0	40.12	9.23 7.764

* 間隙圧から計算算式、つまり実測値にはなる。

** 1ガラス=3.7853 dm³, 1lb (kg)=0.4536kg なので、表にある数字に0.1198をかけるとg/cm⁻³の値になる。

計算は次のとくである。

$$\begin{aligned} \Sigma F &= 5 \times 28 (\text{CH}_2) + 133 (\text{CH}_3) + 3 \times 70 (-\text{O}-) \\ &+ 90 (\text{H}) + 2 \times 440 (\text{ONO}_2) = 1453 \end{aligned}$$

$$V = 224/16 = 14.0$$

よって計算 SP=10.4

以上、SP 値は計算で相当にうまく得られることが分る。なお実測 SP 値はすべて日本化学会編「化学便覧、応用編」(丸善、1973) p.831 からとった。

2. Hoy の場合

この 42 ページにわたる長い論文のうち 35 ページは表である。680 種の化合物について、アルファベット順、SP 値順、沸点(°C)順の 3 構に分けて表 2 のようなデータを掲げている。若者 Hoy はユニオン・カーバイド社の人(エスト・バージニアの Coating Materials 部)と出ているが、大へんな努力の結果の表だらうと圧倒されたことであった。

Hoy は Small の計算法を検討することから始めるのだが、例えばメチル・セロソルブ(2)について。



$\text{C}_2\text{H}_5\text{O}_2$, $M=76$, $d=0.8597$, よって $V=79.19$

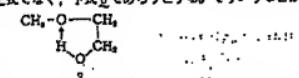
Hoy は臨界圧から算出し SP=11.58 (実測値)とした。

Small によって計算すると、

$$\begin{aligned} \Sigma F &= 214 (\text{CH}_3) + 2 \times 133 (\text{CH}_2) + 2 \times 70 (-\text{O}-) \\ &+ 90 (\text{H}) = 710 \end{aligned}$$

$$\therefore \text{SP} = \Sigma F/V = 8.97$$

全く一致しない。これは例えばメチル・セロソルブは上式でなく、下式でであろうとする。そういうことから



考えなおしたのである。Hoy は新しく表 3 を Molar Attraction Constant として提出した。若干の計算例を出す。

例(1) メチル・セロソルブ

上出のように $V=79.19 \text{ cm}^3 \text{ mol}^{-1}$ である。

表3 Molar Attraction Constant
[(cal/cm³)^{1/4} mol⁻¹] (Hoy)

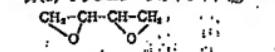
-CH ₃	148.3	CS	209.42
-CH ₂ -	141.5	Cl ₂	342.87
>CH ⁺	85.99	Cl (-O-)	205.08
>C≡	32.03	Cl (二元)	208.27
CH ₃ =	126.54	Cl (芳香族)	161.0
-CH=	121.53	Br	257.88
>CH=	84.51	Br (芳香族)	205.60
-CH= (芳香族)	117.12	C ₆	413.3
-CH= (非芳香族)	98.12	(非芳)	
-O-	114.98	共役族	23.26
O-O- (エボキシ)	176.20	O ₂ 置換	-7.13
-CO-	326.55	O ₂ trans 置換	-18.50
=C=O	282.98	O ₂ 置換	77.6
O-CHO	292.84	C ₆ 置換	20.88
O-COO	567.29	O ₂ ル 置換	6.9
-OH	295.84	O ₂ ル 置換	6.5
-H (後述メターラー)	50.74	O ₂ ラ 置換	40.33
OH (非芳香族)	170.99	O ₂ ラ 置換	22.56
NH ₂	226.56	O ₂ ル 置換	62.5
N-N ⁺	61.08		
C≡N	354.56	基團	
NCO	358.56	基團	135.1

$$\begin{aligned} \Sigma F &= 1483 (\text{CH}_3) + 2 \times 141.5 (\text{CH}_2) \\ &+ 114.98 (-\text{O}-) + 235.84 (\text{OH}) \\ &+ 1351 (\text{基團}) = 807.22 \end{aligned}$$

よって $\text{SP} = \Sigma F/V = 11.46$ と計算される。

これが実測値 11.58 によく一致したと言いうる。

例(2) ブタジエン・ジオキサイド(4)



$\text{C}_4\text{H}_6\text{O}_2$, $M=86$

SP, 実測値 11.78 (Hoyによる)

$d=11.056$, よって $V=77.78$

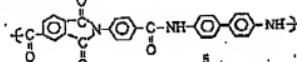
表3による計算。

$$\begin{aligned} \Sigma F &= 2 \times 141.5 (\text{CH}_2) + 2 \times 85.99 (\text{CH}) \\ &+ 2 \times 176.2 (\text{エボキシ}) + 1351 (\text{基團}) \end{aligned}$$

$$\therefore \Sigma F = 942.48$$

よって $\text{SP} = 942.48/77.78 = 12.11$ (実測値² 11.78)

例(3) ポリアミドイミド(5)



$C_{22}H_{14}N_2O_4$, $M=4594$, $d=1.21$ として表3で計算する。それは既に本誌に出ている。文献¹⁾の p.323 の表7 である。それは Ghosh^ら²⁾によるものであったが、

$$\begin{aligned} \Sigma P &= 15(C\text{=O}) + 9(C\text{=芳香環}) \\ &+ 4(C\text{=O}) + 1(-C\text{=O}) + 2(N\text{H}) \\ &+ 1(\text{5員環}) + 1(\text{6員環}) + 1(\text{p-醚}) \\ &+ 1(\text{m-醚}) = 42801 \end{aligned}$$

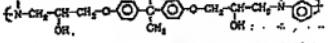
(既になことに該当があった。後の上から5行目の 181.03 は 180.03, 8 行目の 40.03 は 40.33 であった)。そして $\Sigma P = 11.08$ と算出であった。しかし Hoy の表3からみると Ghosh^らは 基礎値の 135.11 を加えるべきであった。即ち

$\Sigma P = (4280 + 135.11) \times 121 / 4594 = 11.47$ となる。が故に (SP 12.10) に届け、N-メチル-2-ピロリドン (SP 11.00) やジメチルホルムアミド (SP 11.79) にいくらか溶けるのなら、今回の SP 11.43 を適用した。私は Ghosh のみを用信せずに、原則によるべきであったと想む。なぜなら Ghosh^らは Hoy の表を用い乍ら Small

の表と寄り合っている。誰しも誤るものだ、と思うと奇妙な気がする。

例(4) 硬化性エポキシ樹脂

もっとも簡単なビスフェノール型エポキシ樹脂を m-



$C_{20}H_{18}O_2N_2$, $M=446$

フェニレンジミンで硬化させたものはまだある。その

ΣP は表3から容易にもとめられる。ただし、

$$\begin{aligned} \text{フェニル} &+ 4(C\text{=O}) + 2(-C\text{=O}) + p \\ &+ 1(\text{m-醚}) \end{aligned}$$

たは m- 増分で計算される。

$$\begin{aligned} \Sigma P &= (2 \times CH_3) + (4 \times CH_2) + (2 \times CH) + (2 \times C) \\ &+ (2 \times O) + (2 \times OH) + (2 \times C_6H_4 - p) \\ &+ (C_6H_4 - m) + 2(C\text{=O}) + (m\text{-増分}) \\ &+ (2 \times N) + (基準値) = 4086.83 \end{aligned}$$

よって硬化物の密度を 1.2 とすると

$$SP = 4086.83 / (446 \div 1.2) = 11.0$$

密度が 1.3 なら SP = 11.9 である。

3. Fedors の場合

Small の場合も Hoy の場合もうまくセットされてい

表4 d_{41} と d_{51} (Fedors, 一部省略)

	d_{41}	d_{51}		d_{41}	d_{51}
CH ₃	1125	32.5	CONH ₂	10000	17.5
CH ₂	1180	16.1	CONH	8000	9.5
CH	820	-1.0	CONC	7050	-7.7
C	350	-19.2	HCON	6600	11.3
H ₂ C=	1030	28.5	HCONH	10500	21.0
-CH=	1030	18.5	COCl	5000	3.0
C=	1030	-5.5	NH ₂	3000	19.2
-C≡	920	27.4	NH	2000	4.5
-C≡C-	1690	8.5	N	1000	-8.0
- (e, m, p)		71.4 ^a	-N≡	2800	5.0
- (e, m, p)		7630 ^a	CN	6100	24.0
- (e, m, p)		33.4 ^a	NO ₂ (油溶性)	7000	24.0
- (e, m, p)		14.4 ^a	NO ₂ (芳香族)	3670 ^a	32.0
- (e, m, p)		-4.6 ^a	NO ₂ (アーティクト)	2800	33.5
(またはフェニル基)			HCO	6800	3.0
-5員環		250	-O-	800	3.8
3, 4員環		750	-OH	7120	10.0
環の中の共役二重結合		400	-CH (ジ炭酸)	5220	13.0
1員環など		-2.2	-S-	3380	12.0
-O-CO-O-		4200	-Cl	2750	24.0
COOH		6600	Cl (ジ炭酸)	2300	26.0
CO		4300	Br	3700	30.0
CO		4150	I	4550	31.5
CHO		5100	Al	3300	2.0
-CO-O-CO-		7300	Si	810	0

(* は略算、本文中のように算出した方がよい)

[557] (54)

ポリマーのSP値の予測一訂正と補足

るが、ポリマーの密度 ρ が必要である。ことに新しい計算法の上で書いて、その SP 値を知らうするとこの ρ が致命的のことになる。その点でこの Fedors の表 4 は役立つことになる。彼はセル容積の V をそれぞれの基のモル容積 d_{V_1} の和と考えるのである。

$$V \approx \sum d_{V_1} \quad (2)$$

そしてその d_{V_1} (25°C) を適当に表 4 のようになる。元来が SP は式(3)である。

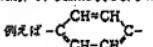
$$SP = (CED)^{1/4} = \frac{(AH - RT)^{1/4}}{V} \quad (3)$$

dH は臨済熱である。この $(AH - RT)$ も、
 $(dH - RT) = \sum d_{E_1}$ (4)

とする。そして

$$SP = \frac{\sum d_{E_1}}{\sum d_{V_1}}^{1/4} \quad (5)$$

とする。表 4 から計算するのだが、+印のフェニル基だけは誤算で、實際は次のように修正した方がよい由。



$$\sum d_{E_1} = 4(\text{CH}_2) + 2(\text{C}) + 3(\text{共役二重結合}) + 6 \text{ 酸素}$$

$$= 4 \times 1030 + 8 \times 1030 + 3 \times 400 + 250 = 7630$$

$$\sum d_{V_1} = 4 \times (135) + 2(-55) + 3 \times (-22) + 16 = 524$$

しかし表中の $d_{E_1} = 7630$, $d_{V_1} = 524$ となっている。

例(1) 簡単なエボキシ樹脂

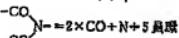
Hoy の場合の例(4)としたままで計算

$$\begin{aligned} \sum d_{E_1} &= 2 \times 1000 (\text{N}) + 4 \times 1180 (\text{CH}_2) + 2 \times 820 (\text{CH}) + 2 \times 7120 (\text{OH}) + 3 \times 7630 (-\text{O}) + 1 \times 350 (\text{C}) + 2 \times 1125 (\text{CH}_3) + 2 \times 800 (-\text{O}-) = 49690 \\ \sum d_{V_1} &= 277 \end{aligned}$$

$$SP = \frac{(49690)^{1/4}}{277} = (17939)^{1/4} = 13.4$$

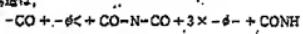
例(2) 上記のポリアミドイミド(5)

までは問題になるのはイミド環だが、次のように考えること。



$$\begin{aligned} \sum d_{E_1} &= 2 \times 4150 + 1000 + 250 = 9500 \\ \sum d_{V_1} &= 2 \times 108 - 9 + 16 = 28.6 \end{aligned}$$

として、あとは足し算すれば足りる。



である。よって上の順に

$$\sum d_{E_1} = 4150 + 7630 + 9500 + 3 \times 7630 + 8000 + 2000$$

$$= 54170$$

$$\sum d_{V_1} = 108 + 334 + 28.6 + 3 \times 524 + 8.5 + 4.5$$

$$= 244$$

$$\left(\frac{\sum d_{E_1}}{\sum d_{V_1}} \right)^{1/4} = \left(\frac{54170}{244} \right)^{1/4} = (222)^{1/4} = 14.9$$

これは大にすぎると思われる。

4. 終りに

以上、前報の誤りを訂正し、加えて新しく全体を紹介した。簡単な便携式ボキシ樹脂の SP 値を算出と言つても結構高分子が溶解するとは思えないが興味がある。以上の 3 種の計算法で、どの方法が最もかは見当が全くつかない。その高分子の密度 ρ が分つておれば Hoy の方法が(それが分らねば Fedors の方法しか仕方がない)何とか計算しやすいように思う。なお Small の方法は實業性さんが 1962 年に紹介されている⁵。なお私は別に⁶ Hoy の手順をまとめておいたが、どういう構造の高分子(または複数)はこれくらいの SP と T_g をもつだらう、と机上の計算をすることができるかは疑いだらう、と思う。

なお、他にも SP (や T_g) の予測法の報告があるだろうと想像する⁷。御教示をたまわることができれば幸甚である。

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US006843819B2

(12) **United States Patent**
Suzuki et al.

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(45) Date of Patent: **Jan. 18, 2005**

(54) AIR FILTER

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§ 371 (c)(1),
(2), (4) Date: **Jan. 23, 2002**

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(51) Int. Cl.⁷ **B01D 46/10**

(52) U.S. Cl. **55/486; 55/385.3; 55/487;**
55/524; 55/528

(58) Field of Search **55/385.3, 486,
55/487, 524, 528, DIG. 5, DIG. 24, DIG. 28;**
428/316.6, 534

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(57) ABSTRACT

There is provided an air filter, which has a long lifetime and permits to capture effectively carbon particles and reduce costs. The air filter has the first filter layer 11 impregnated with oil and the second filter layer 12, which is provided on the downstream side of the first filter layer 11 and serves as a lipophobic layer having an oil-repellent property. The filter material of the first filter layer 11 has a higher density than the filter material of the second filter material 12.

20 Claims, 9 Drawing Sheets

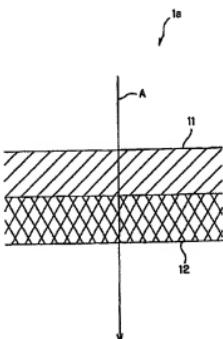


FIG. 1

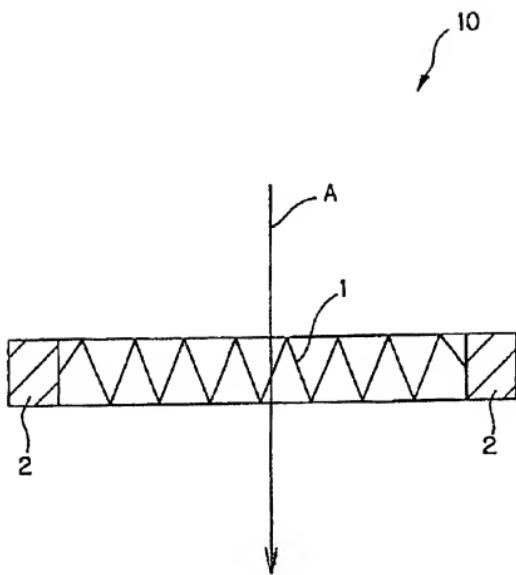


FIG. 2

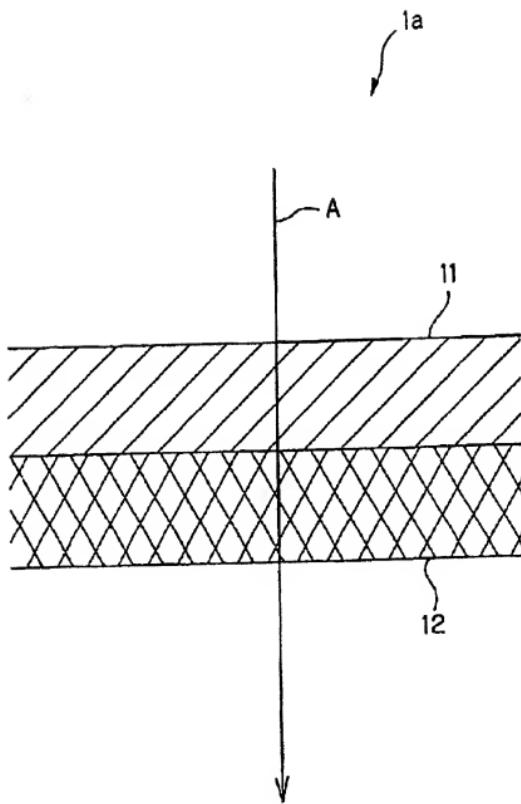


FIG. 3(A)

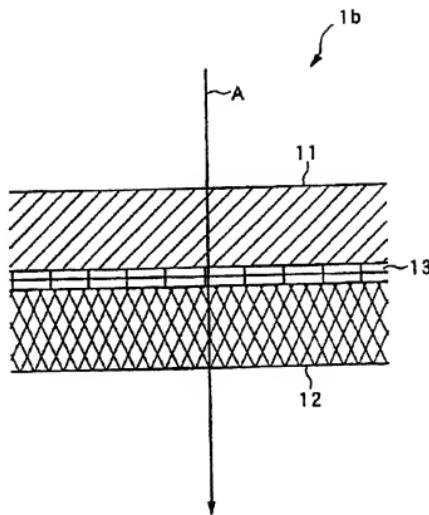


FIG. 3(B)

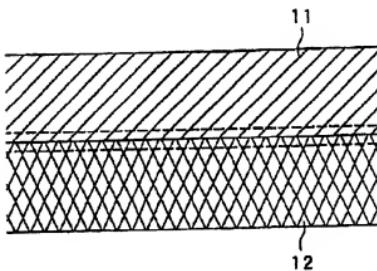


FIG. 4

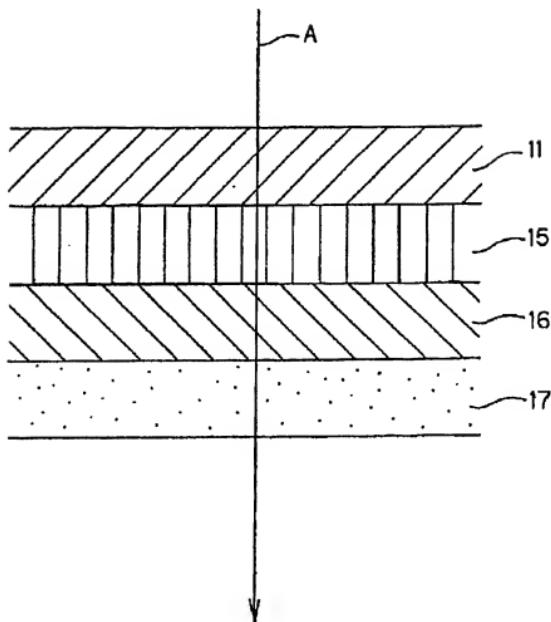


FIG. 5(A)

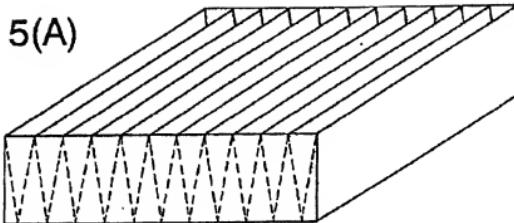


FIG. 5(B)

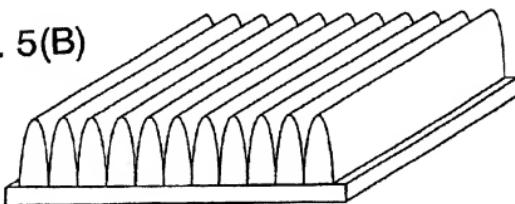


FIG. 5(C)

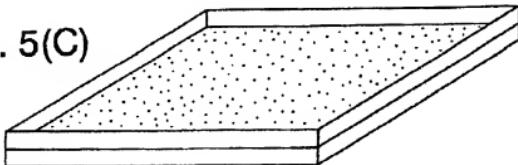


FIG. 6(A)

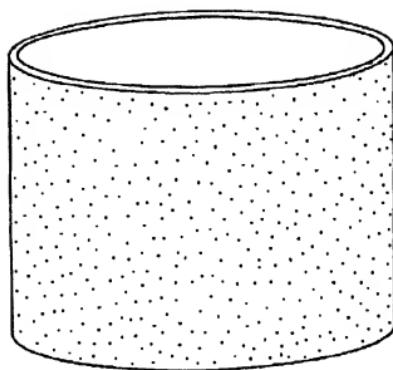


FIG. 6(B)

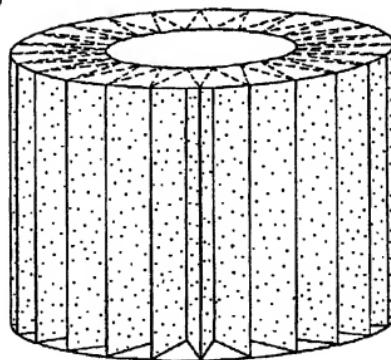


FIG. 7

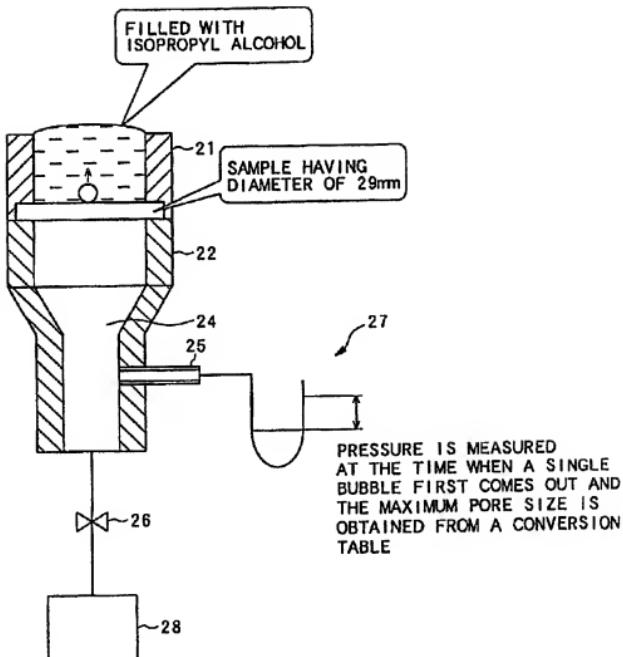


FIG. 8

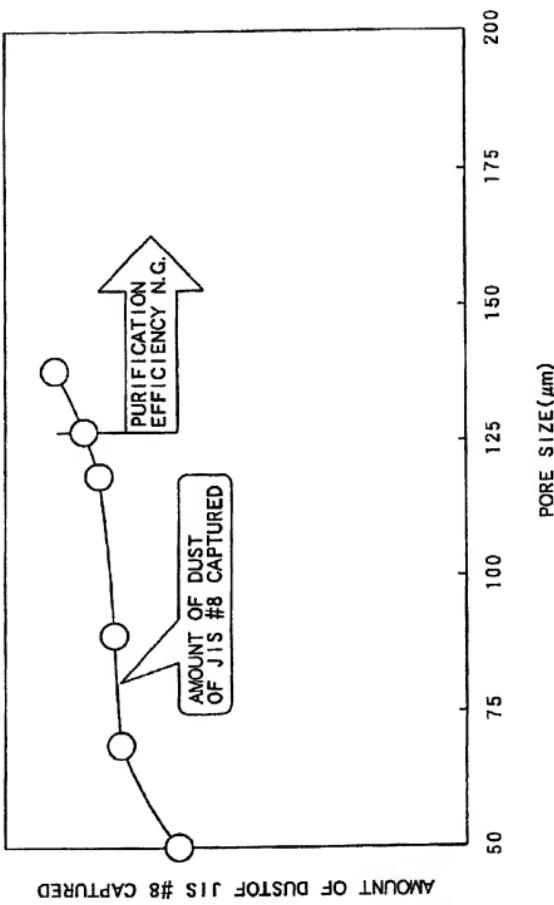
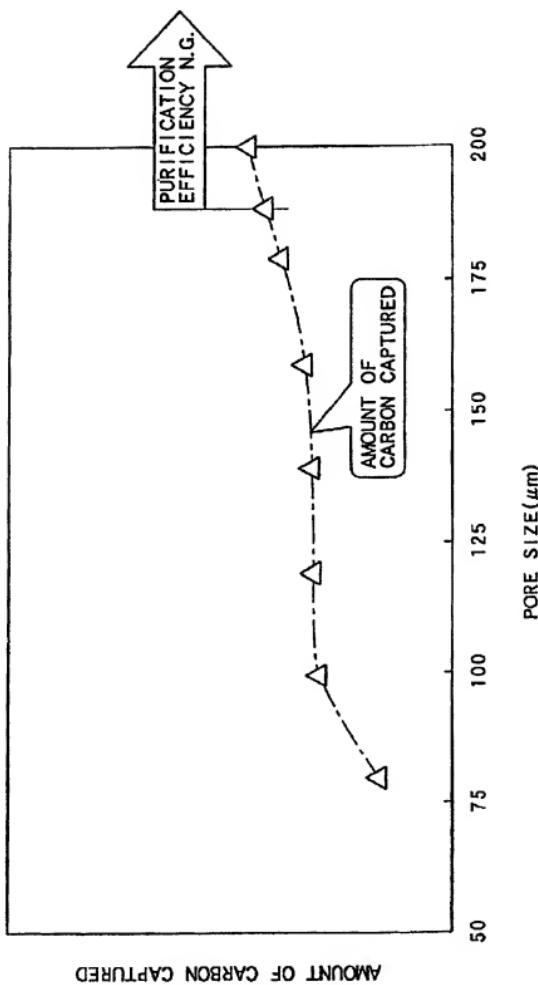


FIG. 9



1
AIR FILTER

TECHNICAL FIELD

The present invention relates to an air filter, and especially to an air filter, which permits to capture efficiently carbon particles.

BACKGROUND OF THE INVENTION

As an air filter for a vehicle, there has been known a wet-type filter having filter paper impregnated with oil. Viscose oil is generally used as impregnation oil. Accordingly, it is called the "viscose type" air filter. The viscose type air filter has an advantage of extended lifetime in comparison with the dry-type filter, which is not impregnated with oil. In the viscose type air filter, oil on the surface of the filter captures dust and then the oil penetrates into the captured dust so as to provide a function of capturing the other dust on the surface of the captured dust. Such a chain capture of dust permits to increase an amount of captured dust per volume of the filter.

Japanese Utility Model Publication No. S63-27767 describes the improved viscose type air filter. The air filter described in Japanese Utility Model Publication No. S63-27767 is a laminate of the first filter material having a relatively high density and the second filter material having a relatively low density. Such a laminate structure permits to make oil impregnation ratios in the first and second filter materials different from each other, resulting in prevention of clogging of the filter, providing an effective capture of dust.

The wet-type filter impregnated with oil has a problem that carbon particles cannot be captured effectively. The dry-type filter, which is not impregnated with oil, can capture effectively carbon, but has a problem of a decreased amount of captured dust per volume of the filter. Use of non-woven fabric leads may cause problems that dust permeation occurs at a high flow rate and in a serious pulsation and costs increase.

DISCLOSURE OF THE INVENTION

An object of the present invention, which was made in view of the above-described circumstances, is therefore to provide an air filter, which permits to provide a large amount of captured dust per volume of the filter, capture effectively carbon particles and reduce costs.

The present invention will be described below. Reference numerals in the accompanying drawings will be given with parentheses in order to facilitate understanding of the present invention. However, the present invention is not limited only to embodiments as illustrated.

In order to attain the aforementioned object, an air filter (10) comprises: a first filter layer (11) impregnated with oil; and a second filter layer (12) provided on a downstream side of said first filter layer, said second filter layer being composed of a lipophobic layer having an oil-repellent property, each of said first filter layer (11) and said second filter layer (12) being formed of filter paper, filter material of said first filter layer (11) having a higher density than filter material of said second filter layer (12), and said second filter layer (12) being composed of said lipophobic layer over an entire thickness.

According to the present invention, the first filter layer that is formed of the filter paper having the high density and is impregnated with oil, captures dust on the one hand, and

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the second filter layer that is formed of the filter paper having the low density and subjected to an oil-repellent treatment so as not to be impregnated with oil, captures carbon particles, on the other hand, thus performing effective filtering operations by means of both filter layers. In general, the thickness of the second filter layer is determined on the basis of specification of capturing the carbon particles. According to the present invention, the second filter layer has the oil-repellent property over the entire thickness.

Consequently, the oil in the first filter layer is not carried by any portion of the second filter layer in its thickness direction, thus making it possible to cause the second filter layer over its entire thickness to capture effectively the carbon particles. On the contrary, in case where the second filter layer is provided only in the upstream portion with a lipophobic layer, the oil in the first filter layer may penetrate through the lipophobic layer, thus causing a problem that the carbon particles cannot be captured effectively over the entire thickness of the second filter layer. Each of the filter layers is formed of filter paper and it is therefore possible to prevent occurrence of dust permeation at a high flow rate and in a serious pulsation and reduce material costs to the minimum. In addition, the filter material of the first filter material has a higher density than the filter material of the second filter layer and pressure loss can also therefore be minimized.

An embodiment of the present invention is characterized in that said first filter layer and said second filter layer are combined integrally with each other.

According to the present invention, the first filter layer and the second filter layer can be combined integrally with each other, thus providing a simple layer structure. The entire thickness of the filter can be adjusted, as an occasion demands.

An embodiment of the present invention is characterized by further comprising an additional layer.

According to the present invention, the additional layer as included can improve performance of the filter element.

An embodiment of the present invention is characterized in that said first filter layer has a pore size of from 70 μm to 120 μm and said second filter layer has a pore size of from 100 μm to 180 μm .

Grounds for limiting the pore size of the first filter layer of from 70 μm to 120 μm and the pore size of the second filter layer of from 100 μm to 180 μm are described below. First, description will be given of grounds for limiting the pore size of the first filter layer of from 70 μm to 120 μm . The present inventors made an experiment on an amount of dust established by the JIS (Japanese Industrial Standard) #8, which penetrates through the first filter, while gradually changing the pore size. FIG. 8 shows the results of the experiment. An abscissa denotes the pore size (μm) and an ordinate shows an amount of captured dust according to the JIS #8, which penetrates the first filter. It is recognized from FIG. 8 that, with the pore size of over 120 μm , a sieve diameter becomes large so as to increase an amount of the dust, which penetrates the first filter, deteriorating purification efficiency. On the contrary, with the pore size of less than 70 μm , an amount of the dust, which penetrates the first filter, is rapidly decreased. Consequently, an amount of the dust captured by the first filter rapidly increases, thus decreasing the lifetime of the filter.

Then, description will be given of grounds for limiting the pore size of the second filter layer of from 100 μm to 180 μm . The present inventors made an experiment on an amount of carbon, which penetrates through the second

filter, while gradually changing the pore size. FIG. 9 shows the results of the experiment. An abscissa denotes the pore size (μm) and an ordinate shows an amount of captured carbon, which penetrates the second filter. It is recognized from FIG. 9 that, with the pore size of over $180 \mu\text{m}$, a sieve diameter becomes large so as to increase an amount of the carbon, which penetrates the second filter, deteriorating purification efficiency. On the contrary, with the pore size of less than $100 \mu\text{m}$, an amount of the carbon, which penetrates the second filter, is rapidly decreased. Consequently, an amount of the dust captured by the first filter rapidly increases, thus decreasing the lifetime of the filter.

Limiting the pore sizes of the first filter layer and the second filter layer within the above-mentioned ranges makes it possible to increase the lifetime of the filter, without deteriorating purification efficiency of the filter.

An embodiment of the present invention is characterized in that said second filter layer has a downstream end, which is exposed.

According to the present invention, it is possible to prevent oil from adhering to the downstream end of the second filter, permitting a more effective capture of the carbon particles utilizing the second filter layer.

The air filter may be manufactured by subjecting the second filter layer to an oil-repellent treatment and then combining the first filter layer and the second filter layer integrally with each other, or by combining the first filter layer and the second filter layer integrally with each other, subjecting the second filter layer to the oil-repellent treatment and then impregnating the first filter layer with oil.

BRIEF DESCRIPTION OF THE DRAWINGS

FIG. 1 is a cross-sectional view illustrating an embodiment of an air filter of the present invention;

FIG. 2 is a cross-sectional view illustrating a filter element of the first embodiment;

FIG. 3 is a cross-sectional view illustrating the filter element of the second embodiment (FIG. 3(A) shows the filter element as being manufactured and FIG. 3(B) shows the filter element as completely manufactured);

FIG. 4 is a cross-sectional view illustrating the filter element of the third embodiment;

FIG. 5 is a view illustrating generally the filter element of the fourth embodiment;

FIG. 6 is a view illustrating generally the filter element of the fifth embodiment;

FIG. 7 is a view illustrating a device for measuring a pore size;

FIG. 8 is a graph illustrating the relationship between the pore size and an amount of dust captured in the first filter layer; and

FIG. 9 is a graph illustrating the relationship between the pore size and an amount of carbon captured in the second filter layer.

BEST MODE FOR CARRYING OUT THE INVENTION

Now, embodiments of an air filter of the present invention will be described below with reference to FIGS. 1 to 6.

An air filter 10 as shown in FIG. 1 is composed of a pleated filter element 1 and a frame member 2, which is formed of plastic into a rectangular shape to support the filter element 1. As shown in FIG. 1, the filter element 1 subjected to an insert injection to form the frame member 2 to which a peripheral portion of the filter element 1 is fixed.

The filter element 1a of the first embodiment as shown in FIG. 2 is provide with the first filter layer 11 having filter element impregnated with oil such as viscose oil and the second filter layer 12, which is composed of a lipophobic layer having an oil-repellent property. Both the first filter layer 11 and the second filter layer 12 are formed of filter paper. The filter material of the first filter layer 11 has a higher density than the filter material of the second filter layer 12. The second filter layer 12 is formed as the lipophobic layer over the entire thickness, i.e., from the upstream-side end face 12a to the downstream-side end face 12b so as to prevent or inhibit oil impregnated in the first filter layer 11 from entering the second filter layer 12 under the function of the oil-repellent property of the lipophobic layer. The downstream-side end face 12b of the second filter layer 12 is exposed to come into contact with air. The upper surface of the first filter layer 11 oxizes with the oil impregnated in the first filter layer 11 to provide a condition in which dust can easily be captured. Imparting the oil-repellent property to the second filter layer 12 makes it possible to provide a filter layer for capturing carbon particles, which have not been captured by the first filter layer 11 impregnated with the oil and passed through the first filter layer 11, under the function of contact filtration.

The lipophobic property can be imparted to the second filter layer 12 with the use of material as properly selected, having the oil-repellent property such as filter paper impregnated with fluorine contained resin. The first filter layer to be impregnated with oil by the subsequent step and the second filter layer, which is formed of filter paper to be impregnated with the above-mentioned fluorine contained resin may be combined together in the paper manufacturing process. It is also possible to provide a combined body of the above-mentioned two layers utilizing a method comprising the steps of combining the first filter layer and the second filter layer together (for example into a filter layer having a single layer structure), impregnating the first filter layer (for example the upper half portion of the above-mentioned single filter layer) with oil and applying lipophobic agent to the second filter layer (for example the lower half portion of the above-mentioned single filter layer). In such a case, an application method such as a spraying method, a roller coating method (such as a direct roller coating method and a kiss-roll coating method), a dipping method and the other method is applicable. Such an application method provides advantages of forming the lipophobic layer with an easy process.

In the second embodiment as shown in FIG. 3(A), the filter element 1a is obtained by combining the first filter layer 11 having the filter material impregnated with oil and the second filter layer 12 composed of the lipophobic layer by an adhesive layer 13. The first filter layer 11 and the second filter layer 12 are joined together so as to permit air to pass through the contacting surfaces of these layers. After completion of the manufacturing steps, the adhesive agent 13 permeates into the first filter layer 11 and the second filter layer 12 so as not to form any gap between the first filter layer 11 and the second filter layer 12 as shown in FIG. 3(B). Also in the embodiment the upper surface of the first filter layer 11 oxizes with the oil impregnated in the first filter layer 11 to provide a condition in which dust can easily be captured. Imparting the oil-repellent property to the second filter layer 12 makes it possible to provide a filter layer for capturing carbon particles under the function of contact filtration. The embodiment describes the state in which no gap is formed between the first filter layer 11 and the second filter layer 12 after completion of the manufacturing steps.

The adhesive layer 13 may however be left between the first filter layer 11 and the second filter layer 12 so that these layers 11 and 12 are apart from each other.

In the method for manufacturing the filter element of the embodiment, the first filter element 11 and the second filter element 12 are placed one upon another through the adhesive agent to combine them together and then the first filter layer 11 is impregnated with oil.

As the adhesive agent used for the adhesive layer 13, it is preferable to select and use, for example a hot-melt material, taking into consideration the manufacturing steps of the filter and conditions under which the filter as manufactured is used. Such a hot-melt material may include olefin material or polyester material.

In the first and second embodiments, an appropriate value may be selected, as the pore size of the first filter layer 11, for example from the range of 70 μm to 120 μm in accordance with a performance required for the air filter. The capturing efficiency of dust in the first filter layer 11 is determined so that the filter sieve can capture the dust. Consequently, the value of the pore size is set taking into consideration the particle size of the dust to be captured and pressure loss.

The pore size of the second filter layer 12 having the lipophobic property is set for example to a value within the range of from 100 μm to 180 μm . Such a setting leads to the pore size of the entire filter element of 70 μm to 120 μm . A reason for using the filter material having a relatively large pore size for the second filter layer 12 is that the second filter layer 12 is to be used to capture carbon particles under the function of contact filtration and it is unnecessary to make the filter material dense. Making the filter material of the second filter layer 12 sparser than the filter material of the first filter layer 11 can rather reduce pressure loss.

The filter element 1c of the third embodiment as shown in FIG. 4 is provided with the first filter layer 11 having the filter material impregnated with oil, the first intermediate layer 15, the second intermediate layer 16 and a clean layer 17, which are disposed in this order from the upstream side to the downstream side. In the embodiment, the second filter layer composed of the lipophobic layer having the oil-repellent property serves as any one of the first intermediate layer 15, the second intermediate layer 16 and the clean layer 17. With respect to the other layer, material, a pore size, thickness and the other conditions may be determined taking into consideration object of use of the filter. The material for the other layer may be formed for example of filter paper or non-woven fabric. In case where the filter paper is used, a water-repellent treatment applied to it makes it possible to prevent water from being sucked into the inside of an engine, even when a suction port of an intake system is located in a place in which water is easily be sucked.

Now, a filtering operation of the air filter of the embodiment of the present invention will be described. An arrow with a reference symbol "A" in FIGS. 1 to 4 denotes a flowing direction of air. As shown in FIGS. 2 to 4, air first passes through the first filter layer 11 and then through the second filter layer.

As described above, the first filter layer 11 is a wet-type filter layer, which is impregnated with oil. A dense structure of the filter material and existence of oil impregnated therein are essential to effective capture of dust. A condition that the filter layer, which is dense and impregnated with oil, exists so as to face an air passage, is also essential to a sufficient capture of dust. Accordingly, use of the first filter layer 11, which satisfies these conditions, makes it possible to effectively capture dust in the air. However, the filter material impregnated with oil cannot generally provide an effective capturing effect of carbon particles in the air. More

specifically, what can effectively be captured by means of the first filter layer 11, is dust other than the carbon particles.

The air, which has passed through the first filter layer 11, enters the second filter layer 12. The second filter layer 12, which is the lipophobic filter having the oil-repellent property and has the sparse (i.e., low density) filter material, can provide an effective capture of the carbon particles. The second filter layer 12 may capture the carbon particles under the function of contact filtration.

According to the air filter of each embodiment of the present invention, the first filter layer 11 effectively captures dust other than carbon particles and the second filter layer 12 effectively captures, over its entirety in the thickness direction, the carbon particles, which have passed through the first filter layer 11. Consequently, the air filter as a whole effectively captures both of the carbon particles and the other dust. In the embodiment of the present invention, the second filter layer 12 captures the carbon particles under the function of contact filtration, unlike the conventional air filter in which the filter materials are disposed so that the density of the filter materials becomes higher in the downstream direction of the flow of air. As a result, it is possible to make the density of the filter material of the second filter layer 12 lower than that of the first filter layer 11, thus reducing pressure loss.

In the wet-type filter, oil generally moves to the downstream side of flow of air through a phenomenon called "carrying off" by air. In the air filter of the embodiment of the present invention, the second filter layer 12 serving as the lipophobic layer is provided on the downstream side of the first filter layer 11. There is ensured a state in which oil does not easily moves to the second filter layer 12 under the function of the oil-repellent property of the lipophobic layer.

In general, an amount of oil impregnated in the filter layer has an influence on the property of capturing carbon particles so that impregnation of the filter layer with oil lower the capturing property. Accordingly, if the second filter layer 12 is impregnated with oil, the carbon particles cannot be captured effectively. In the embodiment of the present invention, the second filter layer 12 is the lipophobic layer as described above so as to prevent the oil from moving. As a result, almost no movement of oil from the first filter layer 11 to the second filter layer 12 occurs. The second filter layer can be kept free from oil, thus making it possible to maintain a high capturing property by which carbon particles can be captured effectively for a long period of time.

In addition, oil does not easily move to the second filter layer 12 and there is no possibility that the oil is carried to the further downstream side of the second filter layer 12. Accordingly, when the air filter of the embodiment of the present invention is mounted on an intake system of an engine for a vehicle, there occurs no inconvenience that the oil is sucked into the downstream side of the air filter. In case where the third filter layer and the fourth filter layer are provided on the downstream side of the second filter layer 12, it is therefore preferable to arrange them so that the densities of these filter layers becomes lower in the flowing direction of air.

In case where the air filters in both embodiments of the present invention are used as an air filter for a vehicle, it is preferable to limit the entire thickness of the filter element 1 within the range of from 0.80 mm to 3.50 mm, and more preferably within the range of from 0.90 mm to 1.75 mm. In such a case, it is preferable to limit the thickness of the first filter layer within the range of from 0.3 mm to 0.75 mm and the second filter layer within the range of from 0.6 mm to 1.00 mm.

In the embodiment of the present invention, the filter element 1 is pleated so as to increase the substantial area of

the filter, as shown in FIGS. 1 to 6. The filter element 1 may be used in an extended flat state as shown in FIG. 5(c). FIGS. 5(a) and 5(b) illustrate examples of the filter elements 1, which are pleated into a panel-shaped filter. It is also possible to form the filter elements 1 into a tubular shape and a chrysanthemum shape as shown in FIGS. 6(a) and (b), respectively, so that air can pass through the filter elements 1 from the inside to the outside, and vice versa. The shape of the filter element can be selected freely in such a manner.

when a single bubble first comes from the upper surface of the filter 23 into the isopropyl alcohol.

A pore size (μm) is obtained on the basis of the thus obtained pressure difference utilizing a conversion table indicated below. The conversion table as conventionally used utilizes a unit "mmAq" for the pressure. A unit "Pa" according to the SI unit system is also indicated in the following table.

TABLE 1

PRESSURE mmAq	PRESSURE Pa	PORE SIZE μm	PRESSURE mmAq	PRESSURE Pa	PORE SIZE μm	PRESSURE mmAq	PRESSURE Pa	PORE SIZE μm
50	490.0	254.3	84	823.2	128.6	118	1156.4	86.0
51	499.8	247.2	85	833.0	126.7	119	1166.2	85.2
52	509.6	240.5	86	842.8	124.9	120	1176.0	84.4
53	519.4	234.1	87	852.6	123.2	121	1185.8	83.6
54	529.2	228.1	88	862.4	121.5	122	1195.6	82.8
55	539.0	222.3	89	872.2	119.8	123	1205.4	82.0
56	548.8	216.9	90	882.0	118.2	124	1215.2	81.3
57	558.6	211.7	91	891.8	116.7	125	1225.0	80.5
58	568.4	206.7	92	901.6	115.2	126	1234.8	79.8
59	578.2	202.0	93	911.4	113.7	127	1244.6	79.1
60	588.0	197.5	94	921.2	112.2	128	1254.4	78.4
61	597.8	193.2	95	931.0	110.8	129	1264.2	77.7
62	607.6	189.0	96	940.8	109.5	130	1274.0	77.0
63	617.4	185.0	97	950.6	108.1	131	1283.8	76.4
64	627.2	181.3	98	960.4	106.8	132	1293.6	75.7
65	637.0	177.6	99	970.2	105.5	133	1303.4	75.1
66	646.8	174.1	100	980.0	104.3	134	1313.2	74.4
67	656.6	170.8	101	989.8	103.1	135	1323.0	73.8
68	666.4	167.5	102	999.6	101.9	136	1332.8	73.2
69	676.2	164.4	103	1009.4	100.7	137	1342.6	72.6
70	686.0	161.4	104	1019.2	99.5	138	1352.4	72.0
71	695.8	158.5	105	1029.0	98.5	139	1362.2	71.4
72	705.6	155.7	106	1038.8	97.4	140	1372.0	70.8
73	715.4	153.0	107	1048.6	96.3	141	1381.8	70.3
74	725.2	150.4	108	1058.4	95.3	142	1391.6	69.7
75	735.0	147.9	109	1068.2	94.3	143	1401.4	69.2
76	744.8	145.5	110	1078.0	93.3	144	1411.2	68.5
77	754.6	143.1	111	1087.8	92.3	145	1421.0	68.1
78	764.4	140.8	112	1097.6	91.4	146	1430.8	67.6
79	774.2	138.5	113	1107.4	90.4	147	1440.6	67.1
80	784.0	136.5	114	1117.2	89.5	148	1450.4	66.6
81	793.8	134.6	115	1127.0	88.6	149	1460.2	66.1
82	803.6	132.4	116	1136.8	87.7	150	1470.0	65.8
83	813.4	130.4	117	1146.6	86.9	151	1479.8	65.1

Brief description will be given below of a testing method 45 for measuring the pore size set forth in the above-described embodiments, with reference to FIG. 7. A disk-shaped filter 23, which serves as an sample and having a diameter of 29 mm, is held air-tightly between two aluminum pipes 21, 22. The lower aluminum pipe 22, which is placed below the filter 23, is provided on the lower end side with a valve 26 for maintaining a constant pressure of air in the pipe 22. The valve 26 is connected to an air supply device 28 having an air compressor.

An air sampler pipe 25 is fitted to the side surface of the aluminum pipe 22. The air sampler pipe 25 is connected to a pressure gauge 27. The pressure gauge 27 can measure the pressure of the inside 24 of the aluminum pipe 22 in this manner.

Prior to measurement of the pore size, the aluminum pipe 21, which is placed above the filter 23, is filled with isopropyl alcohol. The inside 24 of the aluminum pipe 22 is equal to the atmospheric pressure. Then, the valve 26 is gradually opened to supply air from the air supply device 28 so as to gradually increase the pressure of the inside 24 of the aluminum pipe 22. Pressure difference between the atmospheric pressure and the inside 24 of the aluminum pipe 22 is measured utilizing the pressure gauge 27 having a man-

ometer. According to the present invention, the first filter layer, which is formed of a dense filter paper and impregnated with oil, captures dust and the second filter layer, which is formed of a sparse filter paper and is not impregnated with oil, captures carbon particles so that the both filter layers can perform an effective filtration. In addition, the second filter layer has the lipophobic property over its entire thickness. The oil of the first filter layer is not carried to any portion of the second filter layer, thus making it possible for the second filter layer to capture effectively the carbon particles over its entire thickness. Further, the filter material of the second filter layer is sparser than the filter material of first filter layer, thus reducing pressure loss.

What is claimed is:

1. An air filter comprising:
 - a first filter layer formed of a first filter paper material having a predetermined density, said first filter layer being impregnated with oil;
 - a second filter layer provided on a downstream side of said first filter layer so as to be independent from said first filter layer, said second filter layer being formed of a second filter paper material having a lower density than said predetermined density of said first filter layer,

said second filter layer being impregnated over its entirety with an oil-repellent agent.

2. The air filter as claimed in claim 1, wherein: said first filter layer and said second filter layer are combined integrally with each other.

3. The air filter as claimed in claim 1, further comprising an additional layer.

4. The air filter as claimed in claim 1, wherein: said first filter layer has a pore size of from 70 μm to 120 μm and said second filter layer has a pore size of from 100 μm to 180 μm .

5. The air filter as claimed in claim 1, wherein: said second filter layer has a downstream end, which is exposed.

6. The air filter as claimed in claim 2, wherein: said second filter layer is subjected to an oil-repellent treatment and then said first filter layer and said second filter layer are combined integrally with each other.

7. The air filter as claimed in claim 2, wherein: said first filter layer and second filter layer are combined integrally with each other; and then, said second filter layer is subjected to an oil-repellent treatment and said first filter layer is impregnated with oil.

10. The air filter as claimed in claim 2, further comprising an additional layer.

9. The air filter as claimed in claim 3, wherein: said first filter layer has a pore size of from 70 μm to 120 μm and said second filter layer has a pore size of from 100 μm to 180 μm .

10. The air filter as claimed in claim 3, wherein: said second filter layer is subjected to an oil-repellent treatment and then said first filter layer and said second filter layer are combined integrally with each other.

11. The air filter as claimed in claim 3, wherein: said first filter layer and second filter layers are combined integrally with each other; and then, said second filter layer is subjected to an oil-repellent treatment and said first filter layer is impregnated with oil.

12. A air filter, comprising: a first layer of a first filter paper impregnated with oil; and a second layer of a second filter paper placed in a downstream air direction adjacent the first layer,

15. The filter of claim 12, wherein, a downstream face of the second layer is exposed to air, an upper, upstream face of the first layer oozes with the oil.

16. The filter of claim 12, wherein the second layer is impregnated with a resin containing fluorine.

17. The filter of claim 12, further comprising: an adhesive layer binding the first layer with the second layer with the first and second layers contacting one another,

20. The adhesive layer penetrating a downstream side of the first layer and an upstream side of the second layer.

18. The filter of claim 12, wherein the adhesive layer comprises one of an olefin material and a polyester material.

19. The filter of claim 12, wherein, the first layer has a pore size of from 70 μm to 120 μm and the second filter layer has a pore size of from 100 μm to 180 μm .

25. A air filter, comprising: a first layer of a first filter paper impregnated with oil; a second layer of a second filter paper placed in a downstream air direction contacting the first layer; and a hot-melt adhesive layer binding the first layer with the second layer, the adhesive layer penetrating a downstream side of the first layer and an upstream side of the second layer,

26. A density of the first filter paper being greater than a density of the second filter paper,

27. The second layer formed as an oil-repellent lipophobic layer over an entire thickness of the second layer.

28. The filter of claim 12, wherein,

29. The filter of claim 12, further comprising: an adhesive layer binding the first layer with the second layer with the first and second layers contacting one another,

30. The adhesive layer penetrating a downstream side of the first layer and an upstream side of the second layer.

31. The filter of claim 12, wherein the adhesive layer comprises one of an olefin material and a polyester material.

32. The filter of claim 12, wherein, the first layer has a pore size of from 70 μm to 120 μm and the second filter layer has a pore size of from 100 μm to 180 μm .

33. The filter of claim 18, wherein, the second layer is impregnated with a resin containing fluorine.

34. The filter of claim 18, wherein,

35. The filter of claim 18, wherein, the first layer has a pore size of from 70 μm to 120 μm and the second filter layer has a pore size of from 100 μm to 180 μm .

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